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The big idea for all calorimetry processes is energy is conserved. Energy cannot be created or destroyed, but it can be exchanged. $q_{\text{lost}} + q_{\text{gain}} = 0$ or $q_{\text{released}} + q_{\text{gain}} = 0$. Students experience difficulty learning thermochemistry concepts. prove determine the heat exchanged at constant pressure, $q = m c \Delta T$.

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Part 3: Calorimetry shows a basic diagram of a calorimeter, used to measure changes in temperature when combining objects. The Conclusions section contains four bullet points which must be addressed in their conclusion. At this point, I release students to spread out and work.

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